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# **EUROPEAN PATENT APPLICATION**

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⑤④ Improved method of producing amines from alcohols, aldehydes, ketones and mixtures thereof and catalyst used.

⑤⑦ A method of producing amines, the method comprising contacting at reactive conditions at least one alcohol, aldehyde or ketone, or a mixture thereof, with an aminating agent in the presence of a catalyst, is improved by employing as the catalyst a composition comprising cobalt, or nickel, copper and a third component selected from the group consisting of iron, zinc, zirconium and mixtures thereof. For example, polypropylene glycol is quantitatively aminated by contacting it with ammonia at about 175° C and in the presence of a catalyst comprising (metal basis) 30 mole percent cobalt, 63 mole percent copper and about 7 mole percent iron.

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IMPROVED METHOD OF PRODUCING AMINES  
FROM ALCOHOLS, ALDEHYDES, KETONES  
AND MIXTURES THEREOF AND CATALYST USED

5           This invention relates to an improved  
method of producing amines from alcohols, aldehydes,  
ketones and mixtures thereof. In one aspect, the  
invention relates to an ammonolytic method while  
in another aspect it relates to a catalyst useful  
10   therein.

          The prior art describes numerous catalysts  
useful for the ammonolysis of an alcohol, and parti-  
cularly primary and secondary alcohols, to the corres-  
ponding amine. Shirley and Speranza, U.S. Patent  
15   3,128,311, teach a catalyst comprising about 50 to  
90 weight percent nickel, about 10 to 50 weight per-  
cent copper and about 0.5 to 5 weight percent of an  
oxide selected from the class consisting of chromium  
oxide, titanium oxide, thorium oxide, magnesium  
20   oxide, zinc oxide and manganese oxide. Moss, U.S.  
Patent 3,152,998, teaches a catalyst characterized  
by having the composition calculated in mole percent  
on an oxide-free basis of 60 to 85 percent nickel,  
14 to 37 percent copper and 1 to 5 percent chromium.  
25   While these and other catalysts have demonstrated

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ammonolytic utility, catalysts demonstrating superior ammonolytic activity are desirable for reasons of economy, efficiency and convenience.

The present invention is a method for  
5 preparing amines by contacting an alcohol, aldehyde or ketone with an aminating agent in the presence of a catalyst, where the catalyst is composed of nickel or cobalt, copper and iron, zinc or zirconium, characterized in that the  
10 catalyst comprises, calculated in mole percent and on an oxide-free basis:

(a) 20 to 90 percent cobalt and 8 to 72 percent copper, or 20 to 49 percent nickel and 36 to 79 percent copper, and

15 (b) 1 to 16 percent iron, zinc or zirconium.

The present invention also includes a catalyst for preparing amines from at least one alcohol, aldehyde or ketone, or a mixture thereof,  
20 the catalyst consisting of, calculated in mole percent and on an oxide-free basis:

(a) 20 to 90 percent cobalt;

(b) 8 to 72 percent copper; and

(c) 1 to 16 percent of iron, zinc or  
25 zirconium.

The present invention also includes a catalyst for preparing amines from at least one alcohol, aldehyde or ketone, or a mixture thereof,

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the catalyst consisting of, calculated in mole percent and on an oxide-free basis:

- (a) 20 to 49 percent nickel;
- (b) 36 to 79 percent copper; and
- 5 (c) 1 to 16 percent of iron, zinc or zirconium.

These compositions demonstrate superior ammonolytic activity thus allowing quantitative conversion of the alcohol to the corresponding  
10 amine at less rigorous method conditions. Moreover, these compositions demonstrate excellent selectivity and life.

The ammonolytic mechanism of this invention is believed to comprise:

- 15 (a) dehydrogenating the alcohol to the corresponding aldehyde or ketone;
- (b) adding the aminating agent to the aldehyde or ketone to form an imine; and
- (c) hydrogenating the imine to the corresponding amine.
- 20

Consequently, the catalytic compositions of this invention are also useful for the ammonolysis of aldehydes and ketones to the corresponding amines.

The preferred compositions of the present  
25 invention are those comprising, in mole percent and on an oxide-free basis:

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(a) 30 to 80 percent cobalt; 15 to 60 percent copper, and 2 to 14 percent iron, zinc or zirconium, and

(b) 30 to 49 percent nickel; 36 to 69 percent copper, and 1 to 16 percent iron, zinc or zirconium.

The especially preferred compositions are those comprising (same basis):

40 to 70 percent cobalt;  
20 to 50 percent copper, and  
4 to 12 percent iron, zinc or zirconium.

Compositions wherein the third component comprises only zirconium are particularly preferred.

These compositions can be used either unsupported or supported. Typical supports include: alumina, silica, zirconia, zircon (a mixture of zirconia and silica), magnesia, and various cation exchange resins, especially those containing a sulfonated styrene-divinylbenzene copolymer matrix. If the composition is supported, the metal loading (on an oxide-free basis) is usually at least about 0.5 percent and preferably at least about 10 percent of the total weight (support plus composition). The maximum metal content can vary to convenience but it is generally about 50 weight percent and preferably about 20 weight percent.

The catalytic compositions of this invention are readily prepared by any number of different methods but are typically prepared by first precipitating

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the metal components from their salts, e.g., nitrates, chlorides, sulfates, acetates, etc., in a basic, aqueous solution, e.g., sodium or ammonium carbonate; sodium or potassium hydroxide, alkali or alkaline earth metal oxalates, silicates or aluminates, etc. The metal precipitate is then washed, filtered and dried at an elevated temperature, e.g., 60°-180°C, and the dry precipitate is then decomposed at a temperature between about 200°C and about 400°C for a suitable period of time, e.g., 2 hours, to the corresponding oxides. If desired, preparation of the composition can commence with commercially available oxides rather than first preparing the oxides as here described. The resulting oxide mixture is then reduced with hydrogen, sodium borohydride, hydrazine, a reducing metal of greater oxidation potential than cobalt, such as carbon monoxide or some other suitable reducing agent. The degree of reduction is temperature dependent but generally, the first two components (cobalt and copper, or nickel and copper) are reduced to the active metal while the third component, i.e., zinc, iron, zirconium or a mixture thereof, remains an oxide. When this reduction is with hydrogen, a temperature between about 150°C and about 250°C for about 6 to 7 hours is usually adequate. The reduced catalyst is thereafter generally handled in the absence of air. If a supported catalyst is desired, the metal salts can be precipitated directly upon or with the carrier (support).

These compositions can also be prepared from suitable alloys of the three components and at least one leachable fourth component. For example,

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an alloy of either cobalt or nickel, copper, zirconium and aluminum can be formed and subjected to caustic whereby the aluminum is leached from the alloy. The resulting Raney-like structure is then  
5 ready for use.

A catalytic amount of the composition is required for the practice of this invention. The minimum amount of catalyst required will vary with the method reagents and conditions, but a typical  
10 minimum amount is about 0.1 weight percent, and preferably about 1 weight percent, based on the weight of the starting materials. Practical considerations, such as convenience, catalyst recovery and economy, are the only limitations upon the maximum amount of  
15 catalyst that can be used, although these considerations prefer a maximum of about 25 weight percent, and most preferably of about 10 weight percent.

Any alcohol that can be used in known ammonolytic methods can be used in the practice of  
20 this invention. These alcohols comprise a wide variety of hydroxy-containing materials. Representative alcohols include: primary and secondary alcohols, such as alkanols of 1 to about 18 carbon atoms, e.g., methanol, n-propanol, isopropanol,  
25 n-butanol, sec-butanol, isobutanol, the isomeric pentanols and hexanols, 2-ethylhexanol, tridecanol, stearyl alcohol, etc.; cycloalkanols of 5 to 12 carbon atoms, e.g., cyclohexanol, cycloheptanol, etc.; aralkanols of 7 to about 40 carbon atoms,  
30 e.g., benzyl alcohol, 2-phenyl ethanol, etc.; polyhydric alcohols of 2 to about 15 carbon atoms, e.g., ethylene glycol, propylene glycol, 1,3-butanediol,

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1,4-butanediol, 1,5-pentanediol, hexamethylene glycol, decamethylene glycol, 1,12-dihydroxyoctadecane, glycerol, etc.; polymeric polyhydric alcohols, e.g., polyvinyl alcohol; glycol ethers and polyalkylene glycol ethers, e.g., methyl glycol, ethyl glycol, butyl glycol, diethylene glycol, triethylene glycol, tetraethylene glycol, higher polyethylene glycols, dipropylene glycol, tripropylene glycol, polypropylene glycol ether, polybutylene glycol ether, etc.;  
5  
10 aminated alcohols, such as alkanolamines, e.g., ethanolamine, n-propanolamine, isopropanolamine, hexanolamine, diethanolamine, diisopropanolamine, dimethylethanolamine, etc.; and aminated polyhydric alcohols and glycol ethers, e.g., aminated polyethylene glycol, etc. Other suitable hydroxy-  
15  
-containing compounds are disclosed in U.S. Patents 3,347,926; 3,654,370 and 4,014,933.

Any aldehyde or ketone that can be produced from the dehydrogenation of the alcohols here used  
20  
can also be used in the practice of this invention. Representative aldehydes include: methanal, ethanal, propanal, butanal, cyclohexanal, benzylaldehyde, and aldehydes prepared from the dehydrogenation of polyhydric alcohols, polymeric polyhydric alcohols, glycol  
25  
ethers and polyalkylene glycol ethers, aminated alcohols, aminated polyhydric alcohols and glycol ethers. Representative ketones include: propanone, butanone, 2-pentanone, 3-pentanone, 3-methyl-2-butanone, 1-phenyl-2-propanone, acetophenone,  
30  
n-butyrophenone, benzophenone, 3-nitro-4'-methylbenzophenone, and ketones prepared from the dehydrogenation of polyhydric alcohols, polymeric polyhydric alcohols, glycol ethers and polyalkylene glycol ethers, aminated alcohols, aminated polyhydric alcohols and glycol ethers.  
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As used herein, "at least one alcohol, aldehyde or ketone, or a mixture thereof", means that these compounds can be used either singly (which is preferred), or as a mixture of one or more  
5 alcohols, aldehydes or ketones, or a mixture of any two or all three of these types of compounds.

"Alcohol" includes those compounds containing both a hydroxy function and a carbonyl function and "aldehyde" includes those compounds without  
10 a hydroxy function but containing both an aldehyde carbonyl and a ketone carbonyl. Alcohols are preferred to aldehydes and aldehydes are preferred to ketones.

The aminating agents of this invention are  
15 ammonia or primary or secondary amines. The primary and secondary amines generally have alkyl radicals of 1 to about 12 carbon atoms or cycloalkyl radicals of 5 to 8 carbon atoms or aralkyl radicals of 7 to about 40 carbon atoms and include such compounds as:  
20 methylamine, dimethylamine, ethylamine, diethylamine, n-butylamine, sec-butylamine, isobutylamine, ethylenediamine, benzylamine, etc. Other suitable amines include cyclic amines which can contain hetero atoms other than nitrogen, such as oxygen, and these com-  
25 pounds include: morpholine, pyrrolidine, piperidine, piperizine, etc. When ammonia is the aminating agent, primary amines are obtained; when a primary amine is the aminating agent, secondary amines are obtained; when a secondary amine is the aminating agent, ter-  
30 tiary amines are obtained. These aminating agents, like the alcohols, aldehydes and ketones, can also be used either singly (which is preferred) or in

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combination with one another. Primary amines are preferred to secondary amines and ammonia is preferred to primary amines.

5       Stoichiometric amounts of alcohol and  
aminating agent are required for the practice of  
this invention. However, for reasons of convenience  
and efficiency, it is preferable to practice this  
invention with a stoichiometric excess of aminating  
agent to alcohol. Typically, a minimum aminating  
10   agent:alcohol mole ratio of about 1:1 and preferably  
of about 5:1 is employed. Practical considerations,  
such as economy and reaction equipment used, are the  
only limitations upon the maximum ratio, but these  
considerations prefer a mole ratio of about 200:1  
15   and more preferably of about 100:1. The typical  
aminating agent:aldehyde, ketone or mixture mole  
ratios are generally the same as the here recited  
aminating agent:alcohol mole ratios.

20       The method of this invention generally  
employs hydrogen. The amount of hydrogen used can  
vary to convenience, but a typical minimum hydrogen:  
alcohol mole ratio is at least about 0.1:1 and pre-  
ferably at least about 1:1. A typical maximum mole  
ratio is about 50:1 and preferably about 20:1. Again,  
25   the typical hydrogen:aldehyde, ketone or mixture mole  
ratios are generally the same as the here recited  
hydrogen:alcohol mole ratios.

30       Although conventional method conditions  
can here be used, the superior catalytic activity  
of this invention's composition permits the ammonolytic  
process to proceed at lower temperatures and pressures.

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For example, this invention can be practiced at a temperature of at least about 75°C although preferred reaction rates are obtained at a temperature of at least about 120°C. Pressures are of course  
5 dependent upon temperature but a minimum pressure of about 1000 psi (70.5 kg/cm<sup>2</sup>) can be used although a minimum pressure of about 1500 psi (106 kg/cm<sup>2</sup>) is preferred. Again, practical considerations are the only limitations upon the maximum temperature  
10 and pressure, but a maximum temperature of about 400°C and a maximum pressure of about 10,000 psi (705 kg/cm<sup>2</sup>) are preferred. A more preferred maximum temperature is about 250°C and a more preferred maximum pressure is about 6000 psi (423 kg/cm<sup>2</sup>).

15 The invention can be practiced on either a continuous or batch operation, in both the liquid and gas phases, and either neat or in the presence of an inert solvent. By "inert" is meant that the solvent is essentially nonreactive with the method  
20 reagents and products under the conditions used. Exemplary solvents include tertiary alcohols, such as tertiary butanol and tertiary amylalcohol; ethers, such as diethylether and dimethylether; aliphatic and aromatic hydrocarbons, such as hexane, heptane,  
25 cyclohexane, benzene, toluene and xylene; halogenated aliphatic and aromatic hydrocarbons, such as chloroethane, dichloroethane, chlorobenzene, o-dichlorobenzene and chloroform; and nonreactive tertiary amines, such as pyridine, picoline and lutadine.  
30 Moreover, this method can be practiced in either the presence or absence of water, although if water is present, it is preferred that it is not present in amounts greater than about 50 weight percent of the alcohol, aldehyde, ketone or mixture.

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Whether the amines produced by this invention are primary, secondary or tertiary depends not only upon the aminating agent employed, but also upon the particular method conditions used. Short contact time, e.g., between about 0.1 seconds and about 15 minutes, excess ammonia and low temperature and pressure generally favor the production of primary amines. As the aminating agent:alcohol, aldehyde, ketone or mixture mole ratio decreases and/or the temperature increases and/or the contact time increases, secondary and tertiary amines form a larger percentage of the final product. However, longer reaction time favors greater amination of the alcohol. Accordingly, by appropriate selection of aminating agent and process conditions, it is possible to change the proportions of primary, secondary and tertiary amines prepared.

The following examples illustrate the invention. Unless otherwise noted, all parts and percentages are by weight.

#### 20 Catalyst Preparation

The catalysts used in the following examples and controls (unless indicated otherwise) were prepared by adding a solution (0.5 molar) of the corresponding nitrates and a solution (0.5 molar) of ammonium carbonate simultaneously to a reaction flask. The pH (about 6.5) was controlled by the flow rate of the respective solutions. The resulting precipitated carbonates were washed, filtered and subsequently dried at about 60°C. The dry carbonates were then decomposed at about 250°C for about one hour to the corresponding oxides. The oxides were subsequently reduced with hydrogen at a temperature of about 200°C

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for about 6 hours. The reduced catalyst was then handled in the absence of air.

### Procedure

Reduced catalyst (15 g) prepared as above and polypropylene glycol (15 g) of about 400 molecular weight were charged under a protective atmosphere of nitrogen to a rocking autoclave. The autoclave was then charged with ammonia (60 g) and hydrogen (500 psig) and the autoclave contents were subsequently heated for about 5 hours at about 180°C. Afterwards, the autoclave was cooled, vented and the catalyst removed by filtration. The catalyst was then washed with dry methanol and the washings combined with the remaining autoclave contents. Methanol and water were then removed from these contents by a rotary evaporator and the resulting product analyzed either by vapor phase chromatography or titration with hydrochloric acid.

### Examples 1-32

The following examples were performed with catalysts prepared as above and by the above procedure. The figures preceding each catalyst component represent that component's mole percent of the catalyst on an oxide-free basis. The figures reported at the percent amination column represent the weight percent of polypropylene glycol aminated.

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Example	Catalyst	Amination Wt. %
5	1 20-Co, 64-Cu, 16-Zn	26
	2 30-Co, 56-Cu, 14-Zn	36
	3 40-Co, 48-Cu, 12-Zn	29
	4 50-Co, 40-Cu, 10-Zn	27
	*5 60-Co, 32-Cu, 8-Zn	28
10	6 70-Co, 24-Cu, 6-Zn	27
	7 80-Co, 16-Cu, 4-Zn	23
	8 90-Co, 8-Cu, 2-Zn	13
	9 20-Co, 72-Cu, 8-Zr	46
	10 30-Co, 63-Cu, 7-Zr	46
15	11 40-Co, 54-Cu, 6-Zr	52
	12 50-Co, 45-Cu, 5-Zr	53
	13 60-Co, 36-Cu, 4-Zr	47
	14 70-Co, 27-Cu, 3-Zr	45
	15 80-Co, 18-Cu, 2-Zr	45
20	16 90-Co, 9-Cu, 1-Zr	27
	17 20-Co, 64-Cu, 16-Fe	12
	18 30-Co, 56-Cu, 14-Fe	20
	19 40-Co, 48-Cu, 12-Fe	26
	20 50-Co, 40-Cu, 10-Fe	33
25	21 60-Co, 32-Cu, 8-Fe	35
	22 70-Co, 24-Cu, 6-Fe	28
	23 80-Co, 16-Cu, 4-Fe	19
	24 90-Co, 8-Cu, 2-Fe	31
30	25 20-Co, 72-Cu, 8-Fe	54
	26 30-Co, 63-Cu, 7-Fe	60
	27 40-Co, 54-Cu, 6-Fe	52
	28 50-Co, 45-Cu, 5-Fe	55
	29 60-Co, 36-Cu, 4-Fe	51
35	30 70-Co, 27-Cu, 3-Fe	35
	31 80-Co, 18-Cu, 2-Fe	51
	32 90-Co, 9-Cu, 1-Fe	51
	33 30-Ni, 56-Cu, 14-Zn	43.1
	34 ~50-Ni, 45-Cu, 5-Zr	77.6
40	35 30-Ni, 63-Cu, 7-Fe	76.9
	Control	
40	**A 80-Ni, 16-Cu, 4-Zn	11.8
	B 80-Ni, 18-Cu, 2-Zr	29.5
	C 90-Ni, 9-Cu, 1-Fe	60.9

\*Contact time of about 6.25 hours

\*\*Temperature about 170°C

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The control catalysts here used correspond to those known in the art, i.e., catalysts wherein nickel is present in at least about 50 mole percent of the catalyst. A comparison of the example and control data demonstrates the superior activity of the former over the latter.

Example 36

A rocking, 300 cc autoclave was charged with monoethanolamine (20 g), ammonia (75 g), hydrogen at 500 psig (35.3 kg/cm<sup>2</sup>) and a catalytic composition (8 g) comprising cobalt (30 mole percent), copper (63 mole percent) and Fe<sub>2</sub>O<sub>3</sub> (7 mole percent). The charged autoclave was heated to about 180°C and held thereat for 5 hours while continually rocked. Thereafter excess ammonia was vented, the contents filtered and subsequently analyzed by gas chromatography. 67 Percent of the starting monoethanolamine was converted and the product mix contained 43 weight percent ethylenediamine.

Example 37

The same procedure and proportions of components were used as for Example 36, except that the catalyst contained 30 mole percent nickel instead of 30 mole percent cobalt. Conversion of the monoethanolamine was 30 percent and the final product contained 56 weight percent ethylenediamine.

Example 38

A 0.75 inch (1.91 cm) stainless steel tube, one foot long (30.5 cm) was packed with a catalytic composition (66 g) comprising cobalt (50 mole percent), copper (45 mole percent) and ZrO<sub>2</sub> (5 mole percent) and

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was heated by a hot oil bath. The results of several runs at various temperatures are tabulated below. A mixture of monoethanolamine (MEA) and ammonia ( $\text{NH}_3$ :MEA mole ratio of about 7) with a blanket of hydrogen at 500 psig ( $35.3 \text{ kg/cm}^2$ ) was pumped through the packed tube at 1500 psig ( $106 \text{ kg/cm}^2$ ) and 180 cc/hr. Contact time was about 10 minutes.

10 Results of Continuous Ammonolytic Operation with Co-Cu-Zr Composition

<u>Run</u>	<u>Temp. (<math>^{\circ}\text{C}</math>)</u>	<u>Conversion (wt. %)</u>	<u>EDA Yield<sup>1</sup> (wt. %)</u>
A	150	32	82
B	175	78	70
15 C	200	82	40

<sup>1</sup>Ethylenediamine yield based on the weight of the product mix.

Control D

20 Example 38 was repeated except that a commercially available catalytic composition (146 g) comprising cobalt (70-75 mole percent), copper (20-25 mole percent) and chromium oxide (1-5 mole percent) was substituted for the catalytic composition of Example 38. The results are tabulated 25 below.



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Results of Continuous  
Ammonolytic Operation  
with Co-Cu-Cr Composition

5	<u>Run</u>	<u>Temp. (°C)</u>	<u>Conversion (wt. %)</u>	<u>EDA Yield<sup>1</sup> (wt. %)</u>
	A	150	10	76
	B	175	56	52
	C	200	56	--

10 <sup>1</sup>Ethylenediamine yield based on the weight  
of the product mix.

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## CLAIMS:

1. A method for preparing amines by contacting an alcohol, aldehyde or ketone with an aminating agent in the presence of a catalyst, where the catalyst is composed of nickel or cobalt, copper and iron, zinc  
5 or zirconium, characterized in that the catalyst comprises, calculated in mole percent and on an oxide-free basis:

(a) 20 to 90 percent cobalt and 8 to 72 percent copper, or 20 to 49 percent nickel and 36  
10 to 79 percent copper, and

(b) 1 to 16 percent iron, zinc or zirconium.

2. The method of Claim 1 wherein the catalyst comprises, calculated in mole percent  
15 and on an oxide-free basis, about:

(a) 30 to 80 percent cobalt;

(b) 15 to 60 percent copper; and

(c) 2 to about 14 percent of iron, zinc  
or zirconium.

20 3. The method of Claim 1 wherein the catalyst comprises, calculated in mole percent and on an oxide-free basis, about:

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- (a) 40 to 70 percent cobalt;
- (b) 20 to 50 percent copper; and
- (c) 4 to 12 percent of iron, zinc or zirconium.

5           4. The method of Claim 1 wherein the catalyst comprises, calculated in mole percent and on an oxide-free basis, about:

- (a) 30 to 49 percent nickel;
- (b) 36 to 69 percent copper; and
- 10       (c) 1 to 16 percent of iron, zinc or zirconium.

5. The method of Claim 1 wherein the contacting is performed at a temperature of at least 75°C and a pressure of at least 1000 psi (70.5 kg/cm<sup>2</sup>).

15           6. The method of Claim 5 wherein the contacting is performed at a temperature of from 120°-250°C.

7. The method of Claim 5 wherein the contacting is performed under a pressure of from 1000 to 6000 psi (70.5 to 423 kg/cm<sup>2</sup>).

20           8. A catalyst for preparing amines from at least one alcohol, aldehyde or ketone, or a mixture thereof, the catalyst consisting of, calculated in mole percent and on an oxide-free basis:

- (a) 20 to 90 percent cobalt;
- 25       (b) 8 to 72 percent copper; and
- (c) 1 to 16 percent of iron, zinc or zirconium.

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9. The catalyst of Claim 8 consisting of:

- (a) 30 to 80 percent cobalt;
- (b) 15 to 60 percent copper; and
- (c) 2 to 14 percent of iron, zinc or

5 zirconium.

10. The catalyst of Claim 8 consisting of:

- (a) 40 to 70 percent cobalt;
- (b) 20 to 50 percent copper; and
- (c) 4 to 12 percent of iron, zinc or

10 zirconium.

11. A catalyst for preparing amines from at least one alcohol, aldehyde or ketone, or a mixture thereof, the catalyst consisting of, calculated in mole percent and on an oxide-free basis:

- 15 (a) 20 to 49 percent nickel;
  - (b) 36 to 79 percent copper; and
  - (c) 1 to 16 percent of iron, zinc or
- zirconium.



DOCUMENTS CONSIDERED TO BE RELEVANT			CLASSIFICATION OF THE APPLICATION (Int. Cl. <sup>3</sup> )
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	
	<p>GB - A - 936 322 (JEFFERSON CHEMICAL COMP.)</p> <p>* Page 1, lines 24-64 *</p> <p>--</p> <p>CHEMICAL ABSTRACT, vol. 89, no. 21, November 20, 1978, abstract 179532n. Columbus, Ohio, USA</p> <p>&amp; HU - A - 14 908 (PETI NITRO- GENMUVEK) (27-05-1978)</p> <p>--</p>	1-11	C 07 C 85/06 85/08 B 01 J 23/74 23/80
A	<p>GB - A - 1 195 287 (B.A.S.F. A.G.)</p> <p>* Page 1, lines 48-64; examples *</p> <p>--</p>	1	TECHNICAL FIELDS SEARCHED (Int.Cl. <sup>3</sup> )
A	<p>US - A - 4 113 662 (R.G. WALL)</p> <p>* Column 4, line 48 - column 5, line 20 *</p> <p>--</p>	8-10	C 07 C 85/00 B 01 J 23/00
A	<p>CHEMICAL ABSTRACTS, vol. 89, no. 24, December 1978, page 152, abstract 200204u Columbus, Ohio, USA ZHOROV, Yu, M. et al. "Purification of straight-run gasoline from sulfur and nitrogen compounds".</p> <p>&amp; SU - A - 622 836 (GUBKIN, I.M., INSTITUTE OF THE PETROCHEMICAL AND GAS INDUSTRY, MOSCOW)</p> <p>--</p>	11	CATEGORY OF CITED DOCUMENTS
A	<p>US - A - 4 083 799 (J.H. ESTES)</p> <p>* Column 3, lines 50-66; alloy 7 *</p> <p>--</p>	11	X: particularly relevant A: technological background O: non-written disclosure P: intermediate document T: theory or principle underlying the invention E: conflicting application D: document cited in the application L: citation for other reasons
The present search report has been drawn up for all claims			&: member of the same patent family, corresponding document
Place of search The Hague		Date of completion of the search 26-11-1979	Examiner PAUWELS



DOCUMENTS CONSIDERED TO BE RELEVANT			CLASSIFICATION OF THE APPLICATION (Int. Cl. <sup>3</sup> )
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	
A	US - E - 27 296 (YATARO ICHIKAWA) * Example E; table 5: E-2 *  -----	8-10	
			TECHNICAL FIELDS SEARCHED (Int. Cl. <sup>3</sup> )